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WORKSHEET - BALANCING CHEMICAL EQUATIONS

Worksheet: Writing and Balancing Chemical Reactions
1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.
a. $\text{Cu (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{CuO (s)}$
b. $\text{H}_2\text{O (l)} \rightarrow \text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$
c. $\text{Fe (s)} + \text{H}_2\text{O (g)} \rightarrow \text{H}_2 \text{ (g)} + \text{Fe}_3\text{O}_4 \text{ (s)}$
d. $\text{AsCl}_3 \text{ (aq)} + \text{H}_2\text{S (aq)} \rightarrow \text{As}_2\text{S}_3 \text{ (s)} + \text{HCl (aq)}$
e. $\text{CuSO}_4 \cdot 5 \text{ H}_2\text{O (s)} \rightarrow \text{CuSO}_4 \text{ (s)} + \text{H}_2\text{O (l)}$

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UNIT 05 CHEMICAL KINETICS AND EQUILIBRIA NOTES (ANSWERS)

Chemistry AP Unit 5: Chemical Kinetics and Equilibria
Unit 5: CHEMICAL KINETICS AND EQUILIBRIA
Chapter 12: Chemical Kinetics
12.1: Reaction Rates
Chemical Kinetics: - the branch of chemistry that studies the how fast reaction proceeds.
Reaction Rates: - the speed of which the concentration of a reactant or product changes per unit of time. - unit is in mol (L s) or mol (L min)
[A] Concentration A at...

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